

2

AD-A213 186

OFFICE OF NAVAL RESEARCH

TECHNICAL REPORT

FOR

Contract N00014-88K-0369

Technical Report No. 2

Synthesis and Electrochemistry of 2,2'-Bipyridyl
Complexes of Dioxorhenium(V)

by

M. S. Ram, Christopher S. Johnson,
Robert L. Blackburn and Joseph T. Hupp
(312) 491-3504

Department of Chemistry
Northwestern University
Evanston, Illinois 60208-3113

Revised August, 1989

DTIC
ELECTE
OCT 05 1989
S B D

Reproduction in whole, or in part, is permitted for any purpose of the United States Government.

** Accepted for publication in Inorganic Chemistry

DISTRIBUTION STATEMENT A
Approved for public release
Distribution unlimited

89 10 5 003

REPORT DOCUMENTATION PAGE

1a. REPORT SECURITY CLASSIFICATION		1b. RESTRICTIVE MARKINGS	
2a. SECURITY CLASSIFICATION AUTHORITY		3. DISTRIBUTION/AVAILABILITY OF REPORT	
2b. DECLASSIFICATION/DOWNGRADING SCHEDULE			
4. PERFORMING ORGANIZATION REPORT NUMBER(S) 02		5. MONITORING ORGANIZATION REPORT NUMBER(S)	
6a. NAME OF PERFORMING ORGANIZATION Northwestern University	6b. OFFICE SYMBOL (If applicable)	7a. NAME OF MONITORING ORGANIZATION	
6c. ADDRESS (City, State and ZIP Code) Department of Chemistry Evanston, Illinois 60208-3113		7b. ADDRESS (City, State and ZIP Code)	
8a. NAME OF FUNDING/SPONSORING ORGANIZATION Office of Naval Research	8b. OFFICE SYMBOL (If applicable)	9. PROCUREMENT INSTRUMENT IDENTIFICATION NUMBER	
8c. ADDRESS (City, State and ZIP Code) Department of the Navy Arlington, Virginia 22217		10. SOURCE OF FUNDING NOS.	
		PROGRAM ELEMENT NO.	TASK NO.
11. TITLE (Include Security Classification) Synthesis and Electro- Chemistry of 2,2'-Bipyridyl Complexes of ...		PROJECT NO.	WORK UNIT NO.
12. PERSONAL AUTHOR(S) M. S. Ram, Christopher S. Johnson, Robert L. Blackburn and Joseph T. Hupp			
13a. TYPE OF REPORT Technical Report	13b. TIME COVERED FROM _____ TO _____	14. DATE OF REPORT (Yr., Mo., Day) 8/89	15. PAGE COUNT 35
16. SUPPLEMENTARY NOTATION			
17. COSATI CODES		18. SUBJECT TERMS (Continue on reverse if necessary and identify by block number)	
FIELD	GROUP	SUB. GR.	
		multielectron transfer, inorganic electrochemistry	
19. ABSTRACT (Continue on reverse if necessary and identify by block number)			
<p>The rhenium(V) bipyridyl(bpy) complexes [(bpy)(py)₂Re(O)₂]ClO₄ and [(bpy)(3-pySO₃)₂Re(O)₂](Na)·2H₂O have been prepared by reaction of 2,2'-bipyridine with trans-[(py)₄Re(O)₂]Cl and trans-[(3-pySO₃)₄Re(O)₂](Na)₃ in methanol or methanol/water solutions (py = pyridine, 3-pySO₃ = 3-pyridinesulfonate). Despite the seemingly trivial change in coordination environment the bpy-containing complexes were found to behave very differently than the corresponding tetrapyridyl complexes. For example, the pH-independent formal potential for reduction of Re(VI) to Re(V) was lowered by some 600 mV. On the other hand, the pH-dependent potential for the Re(V/III) couple increased by ca. 200 mV upon bpy coordination. The V/III couple also became considerably more reversible (kinetically) in its electrochemical response. The Re(III/II) potential was also shifted in the positive direction. At high pH's, the V/III and III/II couples</p> <p style="text-align: right;">continued on reverse side</p>			
20. DISTRIBUTION/AVAILABILITY OF ABSTRACT		21. ABSTRACT SECURITY CLASSIFICATION	
UNCLASSIFIED/UNLIMITED <input type="checkbox"/> SAME AS RPT. <input type="checkbox"/> DTIC USERS <input type="checkbox"/>			
22a. NAME OF RESPONSIBLE INDIVIDUAL Joseph T. Hupp	22b. TELEPHONE NUMBER (Include Area Code) (312) 491-3504	22c. OFFICE SYMBOL	

coalesced to yield the first example of three-electron redox behavior in rhenium chemistry. By analogy with known dioxo-osmium complexes, both the thermodynamic and kinetic electrochemical effects were attributed to a relative destabilization of the inaccessible Re(IV) oxidation state following bpy coordination. It was proposed that the destabilization is induced by *trans* to *cis* isomerization. The *cis* dioxo configuration has only rarely been seen in d^2 coordination chemistry, and never for hexacoordinate rhenium. Nevertheless, circumstantial evidence from IR, NMR, and Raman experiments supported the assignment. Unfortunately, proof was not possible by these methods since both *cis*- and *trans*-(bpy)(py)₂Re(O)₂⁺ would possess C_{2v} symmetry. Very recent x-ray investigations (to be reported elsewhere) have now confirmed the *cis* assignment.

Abstract:

The rhenium(V) bipyridyl(bpy) complexes $[(bpy)(py)_2Re(O)_2]ClO_4$ and $[(bpy)(3-pySO_3)_2Re(O)_2](Na) \cdot 2H_2O$ have been prepared by reaction of 2,2'-bipyridine with *trans*- $[(py)_4Re(O)_2]Cl$ and *trans*- $[(3-pySO_3)_4Re(O)_2](Na)_3$ in methanol or methanol/water solutions (py = pyridine, 3-pySO₃ = 3-pyridinesulfonate). Despite the seemingly trivial change in coordination environment the bpy-containing complexes were found to behave very differently than the corresponding tetrapyridyl complexes. For example, the pH-independent formal potential for reduction of Re(VI) to Re(V) was lowered by some 600 mV. On the other hand, the pH-dependent potential for the Re(V/III) couple increased by ca. 200 mV upon bpy coordination. The V/III couple also became considerably more reversible (kinetically) in its electrochemical response. The Re(III/II) potential was also shifted in the positive direction. At high pH's, the V/III and III/II couples coalesced to yield the first example of three-electron redox behavior in rhenium chemistry. By analogy with known dioxo-osmium complexes, both the thermodynamic and kinetic electrochemical effects were attributed to a relative destabilization of the inaccessible Re(IV) oxidation state following bpy coordination. It was proposed that the destabilization is induced by *trans* to *cis* isomerization. The *cis* dioxo configuration has only rarely been seen in coordination chemistry, and never for hexacoordinate rhenium. Nevertheless, circumstantial evidence from IR, NMR, and Raman experiments supported the assignment. Unfortunately, proof was not possible by these methods since both *cis*- and *trans*-(bpy)(py)₂Re(O)₂⁺ would possess C_{2v} symmetry. Very recent x-ray investigations (to be reported elsewhere) have now confirmed the *cis* assignment.

By <i>per CB</i>	
Distribution/	
Availability Codes	
Dist	Avail and/or Special
A-1	

Introduction

We and many others have had a long-standing interest in the kinetics and dynamics of single electron transfer (ET) reactions involving metal complexes.¹ Recently that interest has turned towards multi-electron processes. In particular, we have been interested in multi-ET kinetics at electrochemical interfaces. Among the most prominent multi-electron reagents are the many oxo and dioxo complexes of transition metals.² Typically these feature a very rich, yet chemically reversible, redox chemistry.³ Furthermore, they often figure importantly in both artificial and naturally-occurring catalytic systems.²

For our purposes the *trans* dioxo complexes of rhenium (V) were deemed nearly ideal for initial studies.⁴⁻⁶ While not yet particularly noteworthy as redox catalysts, these species do display a pH-dependent, two electron reduction that is evidently kinetically hindered at electrochemical interfaces.⁶ Furthermore the complexes are known to possess interesting photoredox properties.⁵ Nevertheless, before initiating kinetic studies we felt that two minor difficulties - poor aqueous solubility (millimolar or less) and poor stability in lower oxidation states (typically a few hours) - ought to be corrected. We reasoned that both could be accomplished with minimal synthetic effort. As it turns out, the synthetic remedies have been only partially successful. For $(\text{py})_4\text{Re}(\text{O})_2^+$ (py is pyridine), aqueous solubility was readily induced by adding sulfonate substituents. On the other hand, chemical stability was actually worsened by replacement of two monodentate ligands by the chelate 2,2'-bipyridine.

These observations, while useful, would not generally be of sufficient interest to justify separate publication. Much more compelling and surprising,

however, are the rather large redox thermodynamic and kinetic effects induced by the seemingly trivial substitution of 2,2'-bipyridine for a pair of monopyridyl ligands in each of two tetrapyridyl complexes. Specifically, we find that introduction of a single ancillary carbon-carbon bond leads to: 1) very substantial thermodynamic destabilization of Re(V) versus both Re(VI) and Re(III) , and 2) significant kinetic enhancement of the $\text{Re(V)} \rightarrow \text{Re(III)}$ process at the electrode/solution interface. In this report, the first issue and the associated experimental data are examined in some detail; the second is the focus of ongoing experimental work and will eventually be described elsewhere. From the results contained herein the most probable explanation for both observations is that *trans* to *cis* isomerization occurs upon ligand substitution, with resultant destabilization of Re(V) .

One other aspect of the investigation which initially surprised us was the comparative difficulty in synthesizing various bipyridyl complexes of dioxorhenium. In fact we were never able to obtain an authentic *bis* bipyridyl complex. In view of this difficulty we suggest that the single literature report^{4g} on the preparation of $\text{t-(bpy)}_2\text{Re}^{\text{V}}(\text{O})_2^+$ may possibly be mistaken.

Experimental

Materials. Triflic acid, $\text{CF}_3\text{SO}_3\text{H}$, was obtained from Alfa Chemicals. Biogel-P2 was obtained from Biorad laboratories. All other chemicals were obtained from Aldrich or Mallinckrodt Chemicals and were used as received. Aqueous solutions were prepared using water purified by a Millipore system.

Metal Complexes. $\text{ReO}(\text{Cl})_3(\text{triphenylphosphine})_2$ (isomer a) was prepared according to the method of Wilkinson and co-workers.¹⁰

trans-[(py)₄Re(O)₂]ClO₄. This compound was also prepared according to Wilkinson's method, but we found it necessary to add 1% water to the ethanol (or better yet, acetone) reaction mixture. (Water hydrolyses the dimeric oxo-bridged intermediate species.) The perchlorate salt was obtained (always in small quantities) from the chloride by following the procedure of Pipes and Meyer.⁶

trans-[(3-pySO₃)₄Re(O)₂](Na)₃·3H₂O. 50 g of 3-pyridinesulfonic acid (3-pySO₃H) were added slowly with stirring to a suspension of 24 g of NaHCO₃ in 200 ml of water. (More complete neutralization was found to be a complication in the next step; unneutralized 3-pySO₃H did not react with the triphenylphosphine complex.) The solution was rotary evaporated to near dryness at 60°C, and after cooling to room temperature, precipitated further by addition of ethanol. The precipitate was collected by filtration and dried. The salt proved to be completely insoluble in ethanol or acetone, but was very highly soluble in water. About 2 g of ReOCl₃(triphenylphosphine)₂ and 6 g of 3-pySO₃Na were then refluxed in 95% ethanol for 1.5 hrs and the mixture cooled in ice. The precipitated complex was washed thoroughly with acetone. For further reaction with 2,2'-bipyridine, the complex was used in crude form; otherwise, the complex was dissolved in 10 ml water and passed through a Biogel-P2 column (150 ml volume). The central portion of the yellow band was collected into 100 ml of ethanol cooled in ice. The purified product was filtered, washed with ethanol and dried in a vacuum oven at 45°C. Anal. calcd: C, 24.64; H, 2.25; N, 5.75. Found: C, 24.44; H, 2.11; N, 5.54.

[(bpy)(py)₂Re(O)₂]ClO₄. 1 g of [(py)₄Re(O)₂]Cl was stirred at room temperature for 24 hours with 4 g of 2,2'-bipyridine in methanol (70 ml). About 700 ml ether and 50 ml CH₂Cl₂ were added to the solution, which was stirred and

then allowed to settle for 30 min. The precipitated complex was separated by decantation. It was then dissolved in 100 ml water (kept chilled in ice), filtered and loaded onto Sephadex-C-25 cation exchange column (Aldrich). The unreacted yellow tetrapyridine complex was first eluted with 0.1 M NaNO_3 . The ionic-strength was then raised to 0.3 M in NaNO_3 ; the red bipyridyl complex was eluted into a beaker containing 4 ml of saturated NaClO_4 and kept chilled. The precipitated complex was filtered, washed with two 1 ml portions of water, and two 20 ml portions of ether containing 10% acetone. The complex was dried in a vacuum oven at 40°C . Solutions of this complex were always kept cold. Even so, significant decomposition was often evident within 2 hours. It is also worth mentioning that the reaction is reversible: $(\text{py})_4\text{Re}(\text{O})_2^+$ is obtained in good yield following ten minutes of heating in CH_3CN with excess pyridine. Anal. calcd: C, 38.0; H, 2.85; N, 8.86; Cl, 5.13. Found: C, 36.68; H, 2.77; N, 8.59; Cl, 5.60.

$[(\text{bpy})(3\text{-pySO}_3)_2\text{Re}(\text{O})_2](\text{Na})\cdot 2\text{H}_2\text{O}$. 1 g of $[(3\text{-pySO}_3)_4\text{Re}(\text{O})_2]\text{Na}_3$ and 4 g of 2,2'-bipyridine were stirred in 100 ml CH_3OH and 40 ml water. The solution was filtered and stirred for 24 hours. The complex was precipitated by repeated ether + CH_2Cl_2 addition. The complex was finally dissolved in 100 ml water, filtered (while cold), loaded onto a Sephadex-A-25 column and washed thoroughly with ice cold water. The complex was eluted with an ice-cold 0.3 M NaNO_3 solution. The red band was eluted and rapidly rotary evaporated at 35°C to near dryness. The complex, contaminated heavily with NaNO_3 , was precipitated further with an ethanol-ether mixture, filtered and dried. From the crude solid, about 90% of the complex was extracted with CH_3OH and precipitated with ether. The complex was still contaminated significantly with NaNO_3 . It was redissolved therefore, in a minimal amount of water and passed through a

Biogel-P2 column, using cold water as the eluent. The middle fraction of the red band was collected, and rapidly rotary evaporated to near dryness at 35°C. The complex was precipitated with the addition of an ethanol-ether mixture, filtered and dried in a vacuum oven at 40°C. Solutions of the complex were kept cold; otherwise they decomposed significantly in solution within 2 hours. Anal. calcd: C, 32.0; H, 2.66; N, 7.42. Found: C, 31.86; H, 2.42; N, 7.55.

Attempted preparation of $(bpy)_2Re(O)_2^+$. The reaction of $ReOCl_3(triphenylphosphine)_2$ with 2,2'-bipyridine gave a purple-black material, in wet or dry ethanol or acetonitrile in the presence or absence of oxygen. Based on NMR, IR, cyclic voltammetry and elemental analysis, the identity of the product could not be established. Clearly, however, it was not $[(bpy)_2Re(O)_2]X$. The reaction of the same starting materials in either CH_2Cl_2 or DMSO gave red solutions. Workup yielded only an ether-soluble neutral compound, possibly $(bpy)ReOCl_3$. Reactions of 5-Cl-phenanthroline, 5-sulfonato-2,2'-bipyridine (sodium salt)¹¹ and 4,4'-dichloro-2,2'-bipyridine with $ReOCl_3(triphenylphosphine)_2$ in refluxing ethanol gave only unidentifiable blue, purple and green products. The reaction of $(py)_4Re^V(O)_2^+$ and $(3-pySO_3)_4Re^V(O)_2^{3-}$ with 2,2'-bipyridine at $> 40^\circ C$ appeared to give the same products as with the phosphine reaction. Finally, the reaction of 2,2',6,6'-terpyridyl with $(py)_4Re^V(O)_2$ gave a purple material that could not be identified.

Measurements. Background corrected ultraviolet and visible spectra were obtained in a single quartz cell (pathlength = 0.5 cm) with Cary 14 spectrophotometer that had been rebuilt by On-Line Instruments Service. Spectroelectrochemical experiments were carried out, under argon, by using a thin layer cell. The cell contained a gold minigrid working electrode, a

platinum wire counter electrode and a saturated (sodium) calomel reference.⁷ The potential was controlled with a PAR 273 potentiostat. The concentration of the rhenium complex typically was 1 to 2 mM. In acidic solutions, attempts to reduce Re(V) to Re(III) or Re(II) were initially frustrated by concurrent proton reduction. We found, however, that this difficulty could be suppressed by addition of 2.5×10^{-4} M $\text{Pb}(\text{NO}_3)_2$. (Under- as well as overpotentially deposited lead is a well known inhibitor of molecular hydrogen evolution on electrocatalytic surfaces.⁸) In a few instances problems still remained due to homogeneous proton (or water) reduction by electrogenerated Re(II).⁶ Cyclic as well as rotating-disk voltammograms were obtained by using a well polished, ultrasonicated glassy carbon electrode in the usual three-electrode arrangement. Potentials were applied using a Pine Instruments RDE 4 potentiostat. Except as noted, solutions were prepared by dissolving the complexes in appropriate aqueous buffer solutions at $\mu = 0.1$ M. (The buffers or electrolytes were trifluoromethanesulfonic acid and its sodium salt for pH = 1 to 2, chloroacetic acid/chloroacetate for pH = 3, acetic acid/acetate for pH = 4 and 5, phosphate for pH = 7, tris for pH = 8, sodium borate for pH = 9, bicarbonate/carbonate for pH = 10, diethylamine for pH = 11, and sodium hydroxide for pH = 12 and above. Most solutions contained ~ 0.02 M buffer plus ~ 0.08 M sodium trifluoromethanesulfonate.) These were degassed by purging with nitrogen. IR spectra were obtained as KBr pellets in a Mattson FTIR instrument. NMR spectra were obtained in d_6 -acetone or D_2O with a Varian XLA-400 spectrometer. In the spectra the solvent itself was taken as a reference. Elemental analyses were obtained from the Analytical Services Laboratory of Northwestern University or from Galbaith Laboratories. Fast atom bombardment (FAB) measurements were likewise made in NU's Analytical Services Laboratory.

Raman measurements were obtained in aqueous solution (Na_2SO_4 reference) by using a windowless flow cell. The excitation sources and other instrumentation have been described previously.⁹

Results

Vibrational Spectra. Table 1 lists selected IR spectral data for the dioxorhenium(V) complexes. Only the bands in the metal-oxo stretching region ($750\text{--}950\text{ cm}^{-1}$) are given. We note that in $(\text{py})_4\text{Re}(\text{O})_2^+$ an intense band at 820 cm^{-1} has previously been assigned to an asymmetric O-Re-O stretch.⁴¹ This band shifts to 828 cm^{-1} in $(3\text{-pySO}_3)_4\text{Re}^{\text{V}}(\text{O})_2^{3-}$ and 845 cm^{-1} in each of the bipyridyl complexes. (Note that each of these bands is far removed from the Re-O-Re stretch found in μ -oxo rhenium dimeric complexes.^{13,14}) Also seen for the bipyridyl complexes are intense transitions near 904 and 767 cm^{-1} . The higher energy band is most reasonably assigned as the symmetric O-Re-O stretch. The lower presumably is an additional asymmetric stretch, although an assignment as a low frequency bpy mode cannot be ruled out. (Actually, two overlapping bands appear in the 770 to 775 cm^{-1} region.) Interestingly, the same set of bands appears in Raman spectra of the bpy complexes (Table 1). The appearance of all three bands in both spectra clearly indicates a lack of inversion symmetry. (In contrast, in the higher symmetry tetrapyridyl complexes the IR/Raman exclusion principle is followed.) We further observed that of the oxo modes in the bpy complexes only the lowest energy (possibly bpy) and highest energy ones were resonantly enhanced. Finally, we note that the IR of $(\text{bpy})(\text{py})_2\text{Re}(\text{O})_2^+$ in CH_3CN was identical to that obtained in a KBr pellet.

Mass Spectra. A positive-ion FAB spectrum of $(bpy)(py)_2Re(O)_2^+$ contained the expected parent peak at $m/e = 533$. Other large peaks exist at $m/e = 454$ and 375. These can reasonably be assigned as the one and two pyridyl loss fragments. No peaks are observed below the cluster at 375. Also none are found from 535 up to 1200; consequently, dimer formation can be ruled out. Negative-ion FAB on $(bpy)(3-pySO_3)_2Re(O)_2^-$ gave peaks at $m/e = 691$ and below, as expected. (No peaks above 691 were observed.) Finally a total mass determination by osmometry (Galbraith Labs) for $[(bpy)(py)_2Re(O)_2](ClO_4)$ yielded 607g/mole in acetone (633 g/mole expected), further reinforcing the assignment of the bpy complexes as monomers.

1H -NMR Spectra. The proton NMR spectra of the complexes are given in Figure 1. The spectra can be rationalized on the basis of simple AB coupling, although there is one non-first order coupled multiplet in the $(bpy)(3-pySO_3)_2Re^V(O)_2^-$ spectrum. For $(bpy)(py)_2Re(O)_2^+$ (fig. 1b) the spectrum integrates as 2H (doublet, 10.87 ppm), 2H (doublet, 8.70 ppm), 4H (doublet, 8.44 ppm), 2H (triplet, 7.99 ppm), 2H (triplet, 7.88 ppm) and 6H (multiplet, 7.69 ppm). Note that the α protons (pyridine ring) in the bpy complexes are shifted upfield compared to those in the tetrapyridyl species. On the other hand, the α protons of the dipyridyl ligand itself are found relatively far downfield ($\delta > 10$ ppm). Finally we note that the NMR measurements (fig. 2b and d) are very sensitive to trace impurities of the starting tetrapyridyl complexes, as these have 2-4 times the number of equivalent protons at any given resonance. From the extraneous peaks, the purities for both bpy complexes appear to be greater than 96%, at least with respect to starting materials.

UV-vis Spectra of Rhenium(V). The UV-vis spectra of five Re(V) complexes in water (and other solvents) are summarized in Table 2. The bipyridyl complexes display a visible band at about 490 nm with an extinction coefficient of $4.2\text{--}4.6 \times 10^3 \text{ M}^{-1} \text{ cm}^{-1}$. Both spectra are completely independent of pH in the range from 0.5 to 10. (Decomposition occurs below pH 0.5). The visible spectra of the bipyridyl complexes display a solvent dependence that differs from that of the tetrapyridyl species. Most notably, the band near 490 nm in the former moves towards longer wavelength (ca. 525 nm) upon replacement of water by methanol, ethanol, acetonitrile or DMSO. The band near 400 nm also moves to longer wavelength (see Table 2). In the tetrapyridyl complexes, the band at ~445 nm moves to 420 nm, and the band at 331 nm moves to 400 nm, in the same series of solvents (Table 2). Apart from an extra band near 300 nm for the dipyridyl species, the UV spectra for all five complexes are similar.

Visible Spectra of Re(III) and Re(II). Generation of both Re(III) and Re(II) initially proved difficult. For example, chemical reductions with dithionite, hydrazine or Zn(Hg) in acidic aqueous solutions gave color changes, but reoxidation with $\text{Ru}(\text{NH}_3)_6^{3+}$ led to only about 10% recovery of Re(V). Furthermore, at pH > 5, turbidity resulted upon reduction. Bulk electrolysis over a mercury pool also appeared to induce decomposition. In addition, there were difficulties in transferring the reduced solutions to spectral cells. Initial attempts at direct electrolysis (acidic solutions) within a spectral cell proved unsuccessful due to hydrogen and heat evolution at the gold or platinum working electrode, and because of the excessive time required for electrolysis. Eventually we found, however, that a gold minigrid electrode in an optically transparent thin-layer cell was suitable for obtaining spectra of

the reduced bipyridyl complexes. At pH 10, the hydrogen evolution problem was not serious. Because of the extremely slow reduction of Re(V) to Re(III), however, a sufficient potential to generate Re(II) was employed. This species was then reoxidized to Re(III) at -0.9 V. More than 90% of the original Re(V) spectrum could be restored upon applying 0 V. The primary spectral features of the reduced species at pH 10 are given in Table 3 and the full spectrum shown in Figure 2.

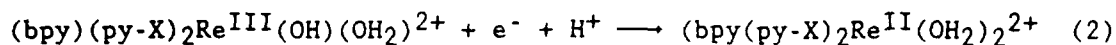
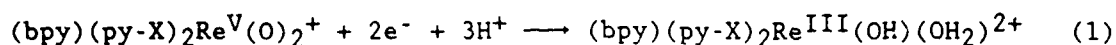
The attempted reduction of the bipyridyl complexes below pH 7 with the gold minigrid was initially unsuccessful. Upon addition of 2.5×10^{-4} M $\text{Pb}(\text{NO}_3)_2$ (see Experimental section) the $\text{Re(V)} \rightarrow \text{Re(III)}$ reduction was possible, but the spectrum of Re(II) could not be obtained. The reversibility of the two-electron transformation was greater than 80% when implemented within an hour of applying the initial reduction potential. The maximum absorption for $(\text{bpy})(\text{py-X})_2\text{Re}^{\text{III}}(\text{OH})(\text{OH}_2)^{\text{n}+}$ was at 383 nm ($\text{X} = \text{H}$) or 450 nm ($\text{X} = \text{SO}_3^-$) (pH 2). For $(\text{bpy})(\text{py-X})_2\text{Re}^{\text{III}}(\text{OH})_2^{\text{n}+}$ λ_{max} is found at 450 nm ($\text{X} = \text{H}$) or 506 nm ($\text{X} = \text{SO}_3^-$) (pH 5 and 10). We are unable to give precise extinction coefficients, but generally they were in the vicinity of $1 \times 10^4 \text{ M}^{-1} \text{ cm}^{-1}$.

Oxidation of Re(V). The dioxorhenium(V) bpy complexes could be oxidized essentially instantaneously in acidic aqueous solution with Ce(IV) , HNO_3 , or Br_2 . Unfortunately, the oxidation product (Re(VI)) was stable for only a few seconds. In neutral solution, rhenium(V) could be oxidized slowly by persulfate, but the product was still unstable. Similarly, Re(VI) instability was observed in spectroelectrochemical experiments. We had hoped that the Re(VI) bipyridyl complexes would be chemically stable because of the large thermodynamic stabilization of Re(VI) relative to Re(V). We note, however,

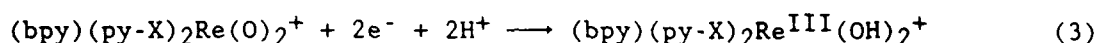
that rhenium(VI) (like Os(VII)) is generally regarded as structurally unstable.^{19,20} Nevertheless, in acetonitrile as solvent Re(VI) was found to be reasonably robust as shown by a successful (slow) bulk electrolysis to Re(VI) and back to Re(V).

Pourbaix Diagrams of Bipyridyl Complexes. Cyclic voltammograms for both the $(py)_4Re^V(O)_2^+$ and $(bpy)(py)_2Re^V(O)_2^+$ species at pH 2 are shown in Figure 3. The overall cyclic voltammetric patterns for all the complexes are similar in that the VI/V, V/III and III/II couples are observed. The limiting currents for these couples in rotating disc experiments, are in the ratio 1:2:1 (figure 4). Note, however, that the VI/V couple in the bpy complex is shifted negative by about 600 mV in comparison to the tetrapyridine. Furthermore, the V/III couple is appreciably more reversible (kinetically) in the bpy case. Latimer diagrams for the various complexes at pH 2 are offered in Figure 5. The pH dependence of the potentials for the various redox couples (Pourbaix diagrams) associated with the bipyridyl complexes are shown in Figures 6 and 7.

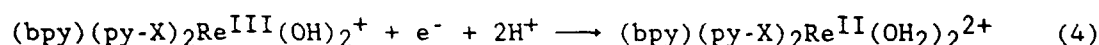
The reduction potentials of the Re(V) bipyridyl complexes are markedly pH dependent, just as for the $(py)_4Re^V(O)_2^+$ complex.⁶ The slopes of $E_{1/2}$ vs pH for the various redox couples are also similar. The potential for the VI/V couple (which is chemically reversible on a CV time scale despite the problems noted above) is independent of pH in the range from 1 to 11. The potentials for the V/III and III/II couples are as follows: In the region $1 < pH < 3.15$, the V/III couple shows a 90 mV/pH slope, while the III/II couple shows a 60 mV/pH slope. These can be represented by equations 1 and 2. (In the following equations $X = H$ or SO_3^- ; for $X = SO_3^-$ the overall charge is 2 units less than indicated.)



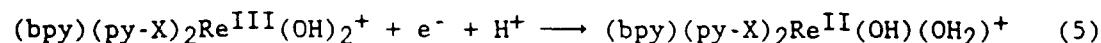
At pH 3.15, the pK_a of $\text{Re}(\text{III})$ is reached. From pH 3.15 to 12 (or higher, depending on X) the Pourbaix diagram for rhenium (V/III) has a 60 mV/slope. In this region the reaction can be represented by eq. 3.



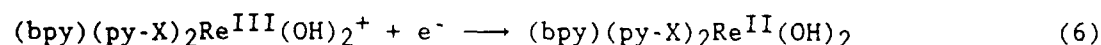
In the pH region 3.15 - 3.75, the rhenium (III/II) couple has a slope of 120 mV. The reaction here can be represented by eq. 4:



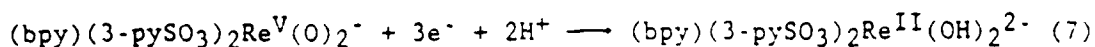
At pH 3.75 the first pK_a of $\text{Re}(\text{II})$ is reached. In the pH region 3.75 - 8.0, the rhenium(III/II) potential varies by 60 mV per pH unit. The reaction here can be described by eq. 5:



At pH 8.0, the second pK_a of $\text{Re}(\text{II})$ is reached. Consequently, in the pH range 8.0 - 13.0 the $\text{Re}(\text{III/II})$ potential is independent of pH. This can be represented by eq. 6:



Finally, for $\text{X} = \text{SO}_3^-$, in the pH range 12 - 13 the V/III and III/II lines merge and the slope becomes 40 mV per pH unit. Furthermore, slow sweep voltammograms within this pH region yield peak separations of less than 28 mV. It follows that the reaction can best be described as:

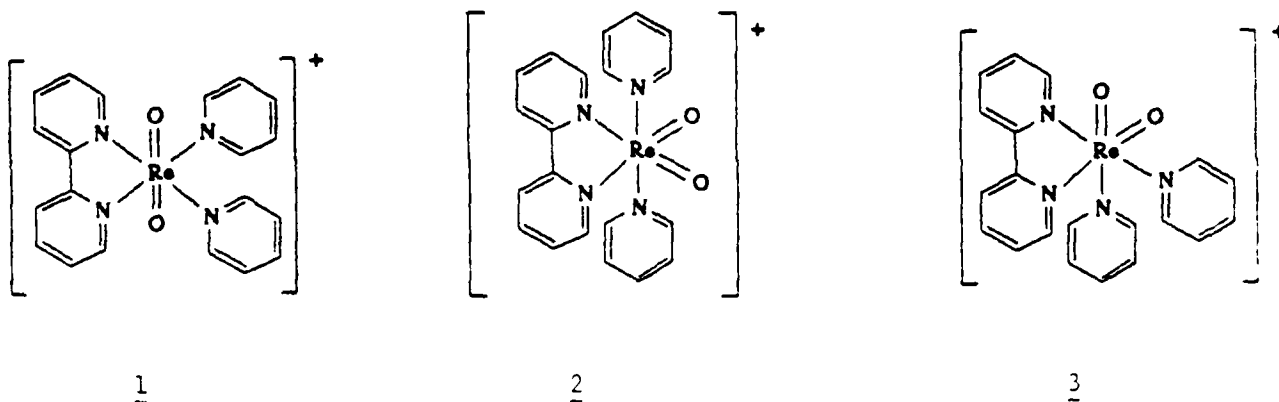


For the corresponding complex with $\text{X} = \text{H}$, the V/III and III/II couples coalesce near $\text{pH} = 13$ (see Figure 6). An extrapolation of the V/III and the III/II lines for the related tetrapyridyl system indicates that a similar three-electron couple will likely exist only at unattainably high pH values.

Discussion

Synthesis. Given the rather striking effects accompanying bipyridyl substitution one must ask what species have actually been synthesized. The sum of NMR, Raman, IR, FAB and elemental analysis studies clearly indicate that the new chelate complexes have the formulae $[(\text{bpy})(\text{py})_2\text{Re}(\text{O})_2](\text{ClO}_4)$ and $[(\text{bpy})(3\text{-pySO}_3)_2\text{Re}(\text{O})_2](\text{Na})$. The isomeric form these take, however, is less clear.

There are three possibilities:



The very low symmetry *cis* isomer (3) can be ruled out on the basis of NMR spectra.¹⁵ The distinction between the *trans* isomer (1) and the alternate *cis* isomer (2) is more problematic. Both possess C_2V symmetry, but lack inversion symmetry. Consequently the selection rules for vibrational spectroscopy within the pair are, strictly speaking, identical. For 1 and 2, symmetric as well as

asymmetric IR and Raman stretches ($O=Re=O$) are permitted. Indeed both types are found in both sets of spectra.

Despite these theoretical considerations, the strictly empirical studies by Griffith on related low symmetry dioxo Os complexes suggest that a spectral distinction between 1 and 2 may still be possible.^{13e,f} Griffith finds for structures like 1 that the pattern of IR and Raman behavior is well explained by approximate D_{4h} symmetry. In other words, intense asymmetric metal-oxo stretches are found in the IR and intense symmetric stretches in Raman spectra, but not vice versa. On that basis we are inclined to designate the bpy-containing species as *cis* dioxo complexes (isomer 2). Additional circumstantial evidence comes from UV-vis and especially electrochemical data (described below). If the assignments are correct the two $Re(V)$ dipyriddy complexes would apparently represent the first examples of solution-stable *cis* dioxo species (hexacoordinate) with d^2 electronic configuration. (Three additional examples are known: $[(bpy)_2Os^{VI}(O)_2](ClO_4)$ ¹⁶ which is stable in the solid state, $[(bpy)_2Ru^{VI}(O)_2](ClO_4)$ which has a transient solution existence¹⁷ and potassium tris(acetato)dioxoosmate, which may or may not be solution stable.²²) Given the rarity of d^2 *cis* complexes, however, it is somewhat surprising that the corresponding *trans*-dioxorhenium(V) complexes (with bpy) have proven elusive.¹⁸

Turning to the last synthesis (or attempted synthesis) we believe that we have obtained the same material described in the literature as $[(bpy)_2Re^V(O)_2](ClO_4)$, although by a different route.^{4h} It is difficult to know for certain because the literature reference lacks UV-vis, IR, Raman and, NMR characterization and analysis for carbon. The blue-purple compound we obtained ($\lambda_{max} = 600 \text{ nm}$ ($\epsilon = 9 \times 10^3 \text{ M}^{-1} \text{ cm}^{-1}$), 571 nm(sh) , and 282 nm ($\epsilon = 3 \times$

$10^4 \text{ M}^{-1} \text{ cm}^{-1}$) changes to purple in 1 M H_2SO_4 ($\lambda_{\text{max}} = 516 \text{ nm}$). The compound has a band in the IR at 873 cm^{-1} , a very poor NMR and a CV featuring only two anodic peaks. Upon reduction with zinc amalgam, the compound decomposes. We do not believe that this is the *cis*-dioxo complex, as the *cis* configuration would likely be unstable for a d^2 species (but note preceding discussion). Furthermore, the IR frequency would probably be lower.^{13,14} It is worth noting that Cotton and co-workers have isolated only $\text{ReCl}_3(\text{bpy})$ from their reactions of chlororhenium complexes¹⁹ and that Johnson, Taha and Wilkinson have reported only an oxo bridged rhenium dimer.²⁰ Lock has pointed out that, unlike $\text{ReOCl}_3(\text{py})_2$ (the first intermediate in the formation of $(\text{py})_4\text{Re}(\text{O})_2^+$ from $\text{ReOCl}_3(\text{triphenylphosphine})_2$), the $\text{ReOCl}_3(\text{bpy})$ species does not hydrolyse the chloride trans to the oxo ligand; instead it forms an oxo bridged dimer.²¹ The above reactivities all imply that our complex is not *trans*-bis(bipyridyl)dioxorhenium(V)⁺. A plausible partial explanation for the apparent inaccessibility of the latter can be found in the probable synthetic intermediacy of *cis*-(bpy)(py)₂Re^V(O)₂⁺ (2).

Electronic Spectra. Even though there is only a slight difference in the gross structure of (bpy)(py)₂Re^V(O)₂⁺ versus (py)₄Re^V(O)₂⁺, tremendous differences exist in both the UV-vis spectra and the electrochemistry. The UV spectrum of either (bpy)(py)₂Re^V(O)₂⁺ or (bpy)(3-pySO₃)₂Re^V(O)₂⁻ (or for that matter, (py)₄Re^V(O)₂⁺ or (3-pySO₃)₄Re^V(O)₂³⁻) can easily be understood as a ligand based ($\pi \rightarrow \pi^*$) transition. (3-pySO₃)₄Re^V(O)₂³⁻ and (py)₄Re^V(O)₂⁺ each display additional bands at 331 and 445 nm. In the bpy complexes these are replaced by bands at 383 and 498 nm. We note, however, that the solvent dependence of the 498 nm band in the bipyridyl complexes differs from that of the tetrapyridyl

complexes (Table 2). This suggests that the transitions have different electronic origins. In the bipyridyl complexes the bands at 383 and 498 nm are tentatively assigned as metal-to-ligand charge transfer transitions. On the other hand, Thorp, Van Houten and Gray have recently assigned the shorter wavelength transition in $(py)_4Re^V(O)_2^+$ as the MLCT, $d_{xy} \rightarrow \pi^*(py)$ transition and the longer wavelength (solvent independent) transition as a d-d transition with intensity stealing from the adjacent MLCT band.^{5c} These differences imply a fundamental difference in electronic structure for the bpy vs. tetrapyridyl complexes. One further indication of substantial structural differences is the complete absence of room temperature luminescence from excited $(bpy)(py)_2Re(O)_2^+$ in a variety of nonaqueous solvents. ($(py)_4Re(O)_2^+$ is highly luminescent.⁵) Evidently in the bpy complex the usually luminescent 3E_g state no longer represents the lowest lying, excited electronic configuration.

For $Re(III)$ and $Re(II)$ Pipes and Meyer have assigned the visible region absorption of the tetrapyridyl complex to $d\pi \rightarrow \pi^*(py)$.⁶ They have pointed out that contributions from other MLCT or LMCT transitions ($p^*(OH) \rightarrow d\pi(Re)$) are also likely. These arguments may well be appropriate for bipyridyl and pyridine sulfonate complexes also.

Electrochemistry. One of the interesting features of dioxo rhenium in comparison to say, osmium¹⁶ or ruthenium,¹⁷ is the tremendous stabilization of the often inaccessible d^1 configuration. For example, in $t-(py)_4Re(O)_2^{n+}$ the d^1 state (Re^{VI}) is attainable at +1.25V, but in analogous Ru and Os species the M^{VII} states are unknown.^{2,16,17} In the rhenium case, the apparent isomerization to the cis form enhances that stabilization, lowering the d^1/d^2 potential to +0.62V. This suggests that the corresponding Re^{VII} and Os^{VII}

might also be observable from *cis*-(bpy)₂M(O)₂⁺, based on oxidation in conventional nonaqueous solvents (i.e. the M(VII/VI) couple might well be accessible at potentials below +2.0V).

For the Re(V/III) couple, bpy substitution and (apparent) *trans* → *cis* isomerization have the opposite effect: E_f is shifted positive in comparison to the tetrapyridyl rhenium couple. The effect can perhaps be partially understood in terms of a relative destabilization of Re^V, as suggested by the shift in VI/V potential (assuming no other states are highly stabilized, or destabilized). A more informative analysis, however, is in terms of d³ stabilization (relative to d²) upon *trans* → *cis* isomerization. The key oxidation state is then the "invisible state" Re^{IV}. Dobson, et al. have noted that in the corresponding (bpy)₂Os^V(O)₂⁺ species, isomerization should lead to a significant change in filled-orbital energetics.¹⁶ In *trans*-dioxo Os the proposed orbital ordering scheme is d_{xy} < d_{xz} < d_{yz}, where the z axis runs along the O=Os=O molecular axis. The ground state configuration for Os^V (or Re^{IV}) will then be (d_{xy})²(d_{xz})¹.

For *cis*-dioxo Re (or Os) the d_{xz} and d_{yz} orbitals are expected to be nearly equivalent.¹⁶ Exact equivalence apparently is absent, judging by the lack of paramagnetic effects in the NMR of Re^V (fig. 1); a probable explanation for the ground-state Re^V diamagnetism can be found in spin-orbit coupling energetics. Under these conditions, Dobson et al.¹⁶ suggest that d_{xz} and d_{yz} are mixed and then slightly split to give dπ₁ < dπ₂ << d_{xy}. This leads to a (dπ₁)², (dπ₂)¹ configuration for either Re^{IV} or Os^V. The important point is that electronic destabilization will be relatively slight in adding one electron to the initially unfilled dπ₂ orbital in *cis* dioxo M; for one electron addition to *trans*-dioxo M the destabilization should be greater.

In the $(\text{bpy})_2\text{Os}(\text{O})_2^{\text{n}+}$ case, Dobson's arguments are amply illustrated by the 370 mV positive shift in the VI/V potential ($\text{pH} = 10$; proton decoupled) upon *trans* \rightarrow *cis* isomerization.¹⁶ In the rhenium case, the corresponding V/IV couple is buried within the overall two-electron, proton-coupled V/III transformation. In other words, at all accessible pH's Re^{IV} will disproportionate into Re^{V} and Re^{III} (or Re^{II}). Nevertheless, the fate of the V/IV couple can still be inferred by noting that $E_{\text{f}}(\text{V/III}) = 1/2 E_{\text{f}}(\text{V/IV}) + 1/2 E_{\text{f}}(\text{IV/III})$. If it is assumed for the moment that the redox transition, $\text{d}^3 + \text{e}^- + \text{nH}^+ \rightarrow \text{d}^4$, is subject to no extra electronic stabilization (or destabilization) upon isomerization, then $E_{\text{f}}(\text{IV/III})(\text{cis}) - E_{\text{f}}(\text{IV/III})(\text{trans})$ will simply equal $E_{\text{f}}(\text{V/IV})(\text{cis}) - E_{\text{f}}(\text{V/IV})(\text{trans})$. The observed *trans* \rightarrow *cis* shift of roughly +200 mV for $\text{Re}(\text{V/III})$ could then be ascribed to an identical shift for the thermodynamically inaccessible V/IV couple. Alternatively, if the change in V/III potential is attributed entirely to the first reduction then $E_{\text{f}}(\text{V/IV})$ would shift by roughly +400 mV, in close agreement with the known $\text{Os}(\text{VI/V})$ behavior.¹⁶ In either case, the shift in $\text{Re}(\text{V/IV})$ potential should also lead to greater $\text{Re}(\text{IV})$ kinetic accessibility (*cis* isomer) and therefore, faster 2e^- reduction kinetics. Faster kinetics indeed are observed for *cis* vs. *trans* (see fig. 3), but the details are not yet fully established.

An overview of the complementary VI/V and V/III potential shifts is provided by the Latimer diagram in fig. 5. From the diagram it can be seen that the hypothetical VI/III reduction potentially is largely unaffected by *trans* \rightarrow *cis* isomerization (i.e. bpy substitution). There is a caveat, however. Over much of the available pH range the *trans* isomer reportedly undergoes only a one proton change.⁶ Reduction of the *cis* isomer, however, always requires at least two protons. Thus the comparison may not be fully appropriate. On the

other hand, the observation of a one-proton change in the tetrapyridyl reaction is surprising since it leads to an apparent violation of the general (empirical) rule against low valent metal oxo formation.^{2b,14} In as much as the assignment of $\text{Re}^{\text{III}}\text{-oxo}$ and $\text{Re}^{\text{II}}\text{-oxo}$ species is predicated upon only slight spectral ($\Delta\lambda_{\text{max}} = 4 \text{ nm}$) and electrochemical changes,⁶ this particular (minor) point might deserve reexamination.

One other interesting aspect of the *cis* $\text{Re}^{\text{V}}(\text{O})_2^+$ chemistry²³ is the appearance of a simultaneous three-electron reduction at $\text{pH} \geq 12$. This is an obvious consequence of: 1) the compression of all of the Re-based redox potentials, and 2) the extended pH dependence of the V/III potential over the range where the III/II couple is pH independent. The only other examples of three electron behavior for monomeric species are apparently those involving reduction of *trans*- $(\text{bpy})_2\text{Os}^{\text{VI}}(\text{O})_2^{2+}$ (low pH's),¹⁶ $(\text{terpyridine})\text{Os}^{\text{VI}}(\text{O})_2(\text{OH}_2)^{2+}$ (low and intermediate pH's)^{3a} and $(\text{phenanthroline})\text{Os}^{\text{VI}}(\text{O})_2(\text{OH})_2$ (high pH's).^{3a} One of the striking aspects of the $\text{Re}(\text{V/II})$ electrochemical reaction, in comparison to the osmium examples, is just how facile the rhenium reaction is kinetically.

Acknowledgements.

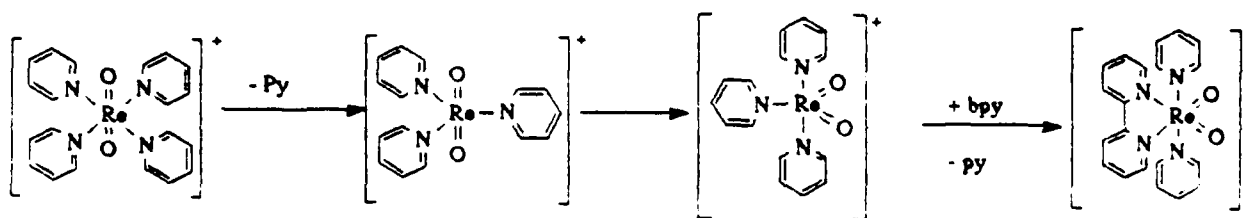
This work was supported by a grant from the Office of Naval Research. The Raman facilities are part of the Materials Research Center at Northwestern (NSF-DMR-8821571).

References

1. Representative reviews: (a) Sutin, N. *Prog. Inorg. Chem.* 1983, 30, 441.
(b) Weaver, M. J., *Comprehensive Chemical Kinetics*, Vol. 27, R. G. Compton, Editor, Elsevier, Amsterdam, 1988. (c) Haim, A. *Comments Inorg. Chem.* 1985, 4, 113. (d) Cannon, R. D., "Electron Transfer Reactions", Butterworths, London, 1980. (e) Wilkins, R. G., "The Study of Kinetics and Mechanisms of Reactions of Transition Metal Complexes", Allyn and Bacon, Boston, 1974. (f) Sutin, N. *Acc. Chem. Res.* 1968, 1, 225.
2. Recent reviews: (a) Meyer, T. J., *J. Electrochem. Soc.* 1984, 131, 221c.
(b) Holm, R. H. *Chem. Rev.* 1987, 1401.
3. Selected references: (a) Pipes, D. W.; Meyer, T. J. *Inorg. Chem.* 1986, 25, 4042. (b) Che, C. M.; Wong, K. Y.; Leung, W. H.; Poon, C. K. *Inorg. Chem.* 1986, 25, 1809. (c) Groves, J. T.; Ahn, K. H. *Inorg. Chem.* 1987, 26, 4391. (d) Marmion, M. E.; Takeuchi, K. J. *J. Am. Chem. Soc.* 1988, 110, 1472. (e) Charney, L. M.; Finklea, H. O.; Schultz, F. A. *Inorg. Chem.* 1982, 21, 549.
4. (a) Beard, J. H.; Calhoun, C.; Casey, J.; Murmann, R. K. *J. Am. Chem. Soc.* 1968, 90, 3389. (b) Lock, C. J. L.; Turner, G. *Acta Crystallogr., Sect. B* 1978, B34, 923. (c) Beard, J. H.; Casey, J.; Murmann, R. K. *Inorg. Chem.* 1971, 10, 2353. (d) Murmann, R. K. *Inorg. Synth.* 1968, 8, 173. (e) Murmann, R. K.; Schlemper, E. O. *Inorg. Chem.* 1971, 10, 2352. (f) Fenn, R. H.; Graham, A. J. *J. Chem. Soc., A* 1971, 2880. (g) Chakravorti, M. C. *Inorg. Nucl. Chem.* 1975, 37, 1991. (h) Chakravorti, M. C. *J. Ind. Chem. Soc.* 1967, 44, 809. (i) Beard, J. H.; Casey, J.; Murmann, R. K. *Inorg. Chem.* 1965, 4, 797.

5. (a) Winkler, J. R.; Gray, H. B. *J. Am. Chem. Soc.* **1983**, *105*, 1373. (b) Winkler, J. R.; Gray, H. B. *Inorg. Chem.* **1985**, *24*, 346. (c) Thorp, H. H.; Van Houten, J.; Gray, H. B. *Inorg. Chem.* **1989**, *28*, 889.
6. Pipes, D. W.; Meyer, T. J. *Inorg. Chem.* **1986**, *25*, 3256.
7. Heineman, W. R. *J. Chem. Ed.* **1983**, *60*, 305.
8. Adzic, R. R. *Isr. J. Chem.* **1979**, *18*, 166.
9. Doorn, S. K.; Hupp, J. T. *J. Am. Chem. Soc.*, **1989**, 111, 4704.
10. (a) Johnson, N. P.; Lock, C. J. L.; Wilkinson, G. *Inorg. Synth.* **1967**, *9*, 145. (b) Johnson, N. P.; Lock, C. J. L.; Wilkinson, G. *J. Chem. Soc.* **1964**, 1054. (c) Johnson, N. P.; Taha, F.; Wilkinson, G. *J. Chem. Soc.* **1964**, 2614.
11. Anderson, S.; Constable, E. C.; Seddon, K. R.; Turp, J. E.; Baggott, J. E.; Pilling, M. J. *J. Chem. Soc., Dalton Trans.* **1985**, *22*, 47.
12. Maerker, G.; Case, F. H. *J. Am. Chem. Soc.* **1958**, *80*, 2745.
13. (a) Griffith, W. P. *J. Chem. Soc., A* **1969**, 211. (b) Griffith, W. P.; Wickens, T. D. *J. Chem. Soc., A* **1968**, 400. (c) Griffith, W. P.; Hawkin, D. J. *J. Chem. Soc., A* **1966**, 472. (d) Griffith, W. P. *Coord. Chem. Rev.* **1970**, *5*, 459. (e) Collin, R. J.; Jones, J.; Griffith, W. P. *J. Chem. Soc., Dalton Trans.* **1974**, 1094. (f) Schroeder, M.; Nielson, A. J.; Griffith, W. P. *J. Chem. Soc., Dalton Trans.* **1979**, 1607. (g) Freni, M.; Girsto, D.; Romiti, P.; Minghetti, G. *Gazz. Chim. Ital.*, **1969**, *99*, 286.
14. Nugent, W. A.; Mayer, J. M., "Metal-Ligand Multiple Bonds", John Wiley and Sons, New York, 1988, pp. 22, 36.
15. For illustrations of this strategy see: Durham, B.; Walsh, J. L.; Carter, C. L.; Meyer, T. J. *Inorg. Chem.* **1980**, *19*, 860.

16. Dobson, J. C.; Takeuchi, K. J.; Pipes, D. W.; Geselowitz, D. A.; Meyer, T. J. *Inorg. Chem.* 1986, 25, 2357.
17. Dobson, J. C.; Meyer, T. J. *Inorg. Chem.*, 1988, 27, 3283.
18. A dissociative mechanism, as sketched below, might provide a kinetic basis for favoring the *cis* geometry. Presumably the second step in the mechanism would be driven by the need to relieve steric crowding in the vicinity of the pyridine α protons. Based on this speculation, the eventual synthesis of a *trans* dioxo rhenium bipyridyl complex would probably require a much more reactive starting material or intermediate species than employed here (so that substitution could outrun isomerization).



19. (a) Cotton, F. A.; Oldham, C.; Walton, R. A. *Inorg. Chem.* 1967, 6, 214.
(b) Jaecker, J. A.; Murtha, D. P.; Walton, R. A. *Inorg. Chem. Acta* 1975, 13, 21.
20. Johnson, N. P.; Taha, F.; Wilkinson, G. J. *Chem. Soc.* 1964, 2614.
21. Guest, A.; Lock, C. J. L. *Canad. J. Chem. Soc.* 1971, 49, 603.
22. Behling, T.; Capparelli, M. V.; Skapski, A. C.; Wilkinson, G., *Polyhedron*, 1982, 1, 840.
23. Based upon an x-ray crystallographic investigation, we have now confirmed (some months after the original manuscript submission) that the $(bpy)(py)_2Re(O)_2^+$ complex indeed does possess a *cis* configuration. The complete details of this study will be reported elsewhere.

Captions for Figures

Figure 1. ^1H -NMR of dioxorhenium complexes (reference solvent). a.

$(\text{py})_4\text{Re}(\text{O})_2^+$ complex in acetone- d_6 δ 9.25(4, 8H), 7.94(t, 4H),

7.68(6, 8H). b. $(\text{bpy})(\text{py})_2\text{Re}(\text{O})_2^+$ complex in acetone- d_6 δ 10.87(d,

2H), 8.70(d, 2H), 8.44(d, 4H), 7.99(t, 2H), 7.88(6, 2H), 7.69(m, 6H).

c. $(3\text{-pySO}_3)_4\text{Re}(\text{O})_2^{3-}$ complex in D_2O δ 9.28(s, 4H), 9.11(d, 4H),

8.26(d, 4H), 7.77(t, 4H). d. $(\text{bpy})(3\text{-pySO}_3)_2\text{Re}(\text{O})_2^-$ complex in D_2O δ

10.36(d, 2H), 8.31(m, 4H), 8.12(d, 2H), 7.4(d, 2H), 7.74(m, 4H),

7.59(t, 2H).

Figure 2. Spectral changes accompanying the reduction of $(\text{bpy})(\text{py})_2\text{Re}(\text{O})_2^+$ and $(\text{bpy})(3\text{-pySO}_3)_2\text{Re}(\text{O})_2^-$ complexes to $\text{Re}(\text{III})$ and $\text{Re}(\text{II})$ at pH 10 in thin layer cell with a gold minigrid working electrode.

Figure 3. Cyclic voltammogram of the $(\text{bpy})(\text{py})_2\text{Re}(\text{O})_2^+$ and $(\text{py})_4\text{Re}(\text{O})_2^+$ complexes at pH 2, using a glassy carbon working electrode.

Figure 4. RDE limiting currents for the various redox processes involving $(\text{bpy})(\text{py})_2\text{Re}(\text{O})_2^+$ at pH 6.8. Rotation rate = 1000 rpm, scan rate = 5 mV/s, $[\text{Re}^V] = 1.67 \text{ mM}$.

Figure 5. Latimer diagram for dioxorhenium complexes at pH 2.

Figure 6. Pourbaix diagram for the $(\text{bpy})(\text{py})_2\text{Re}(\text{O})_2^+$ complex.

Figure 7. Pourbaix diagram for the $(\text{bpy})(3\text{-pySO}_3)_2\text{Re}(\text{O})_2^-$ complex.

Table 1. Infrared and Raman Spectra (750-950 cm^{-1})

Dioxorhenium (V) Complexes

Complex	IR ^a	Raman ^b
$[(\text{bpy})(\text{py})_2\text{Re}(\text{O})_2](\text{ClO}_4)$	767,845,904 cm^{-1}	770,845,905 cm^{-1}
$[(\text{bpy})(3\text{-pySO}_3)_2\text{Re}(\text{O})_2](\text{Na})$	767,845,904	
$[(3\text{-pySO}_3)_4\text{Re}(\text{O})_2](\text{Na})_3$	769,828	
$[(\text{py})_4\text{Re}(\text{O})_2](\text{ClO}_4)$	820	916 ^c

a. KBr pellet. (The first complex was also examined in CH_3CN and gave identical results.) b. aqueous solution. c. ref. 5c.

Table 2 UV-vis Spectra of Dioxorhenium(V) Complexes

Re(V) Complex	Solvent	$\lambda_{\max}(\epsilon)$ nm ($M^{-1}cm^{-1}$)			
$(py)_4Re(O)_2^+$	water	445(1.2×10^3)	331(1.94×10^4)		245(1.53×10^4)
	CH ₃ OH	429	344		
	CH ₃ CH ₂ OH	416	347		
	CH ₃ CN	422	352		
	DMSO	423	356		
$(4\text{-picoline})_4Re(O)_2^+$	water	453(1.24×10^3)	331(2.2×10^4)		245(1.37×10^4)
$(bpy)(py)_2Re(O)_2^+$	water	498(4.2×10^3)	383(8.3×10^3)	300(1.64×10^4)	245(1.45×10^4)
	CH ₃ OH	521	400		
	CH ₃ CH ₂ OH	523	402		
	CH ₃ CN	525	406		
	DMSO	527	413		
$(3\text{-pySO}_3)_4Re(O)_2^{3-}$	water	433(1.4×10^3)	346(2.1×10^4)		253(1.9×10^4)
	CH ₃ OH	419	354		
	DMSO	434	361		
$(bpy)(3\text{-pySO}_3)_2Re(O)_2^-$	water	490(4.6×10^3)	411(9.7×10^3)	300(1.64×10^4)	254(1.4×10^4)
	CH ₃ OH	516	418		
	CH ₃ CH ₂ OH	519	417		
	CH ₃ CN	521	413		
	DMSO	526	421		

Table 3. Visible Spectra of Rhenium (III) and Rhenium (II) at pH 10

Re(V) Complex	Re(III) $\lambda_{\max}(\epsilon)$ nm ($M^{-1} \text{ cm}^{-1}$)	Re(II) $\lambda_{\max}(\epsilon)$ nm ($M^{-1} \text{ cm}^{-1}$)
(py) ₄ Re(O) ₂ ⁺	409 (1.1×10^4)	466
(bpy)(py) ₂ Re(O) ₂ ⁺	454 (1.0×10^4)	501 (1.3×10^4)
(3-pySO ₃) ₄ Re(O) ₂ ³⁻	494	not measured
(bpy)(3-pySO ₃) ₂ Re(O) ₂ ⁻	506 (9×10^3)	546 (1.3×10^4)

